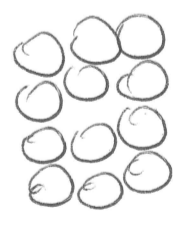


• regular formation
 • strong electrostatic forces in all directions
 • (+) attracted to (-)
 • cannot conduct (no delocalised electrons)



GIANT LATTICE

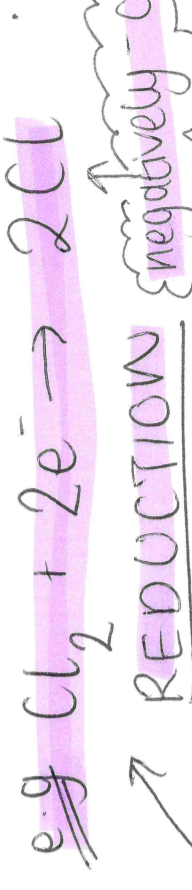
high mpts & bpts
 soluble
 often white solids when electrons are transferred between atoms to get a full outer shell

lower down grp 7 = reactivity decreases = harder to GAIN e⁻

lower down grp 1 = reactivity increases = more likely to lose e⁻

lose e⁻ Grps 1, 2, 6 & 7 are most likely to form ions.

FORMATION



IONIS

Electrolysis

OXIDATION AT ANODE

Aqueous = OH⁻ & H⁺ (and metal ions)

(-) ions go to the anode (anions)

(+) ions go to the cathode (cations)

Molten = only metal ions

PANIC

spills up ionic compounds

REDUCTION AT CATHODE

* Barium chloride & solution you think has sulfate ions in it
 → (+) Barium sulfate white precipitate

* Silver Nitrate + halogen
 → Silver halide (insoluble)

Solubility in H₂O

Ionic compounds dissolve in water

OXIDATION (Loss of e⁻)

REDUCTION (Gain of e⁻)

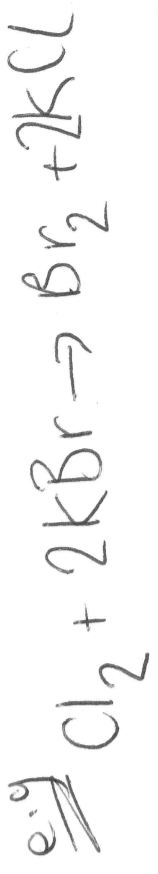
METALS → e.g.



the metal has displaced the "H" from the HCl to make a metal salt & Hydrogen

DISPLACEMENT RXNS

GRP 7's A more reactive grp 7 (e.g. near the top, for Cl) displaces a less reactive grp 7.



Strong acids fully ionise in water
 Weak acids only partially dissociate & give off a smaller proportion of H⁺ ions.